8. Complete the table below.

Element			Ion			
Element Symbol	**Element Name**	**At #**	**Change in Number of e⁻**	**Ion Symbol**	**Name of Ion**	**Noble Gas With Same Number of Electrons**
Cl	chlorine	17	gains 1	Cl^-	chloride	argon
		13	loses 3			
O						
		20				
					sulfide	
			loses 1			argon
				H^+		nil
H					hydride	helium
Mg						
		3				
			gains 1			neon

SCIENCE 90
NOMENCLATURE PRACTISE

PART A. BINARY IONIC COMPOUNDS

1. Name the following compounds.

a. SrO b. Al_2S_3 c. HBr

d. $AlCl_3$ e. MgO f. CsF

g. Ca_3P_2 h. MgS i. CaO

j. BaF_2 k. K_2O l. $BaCl_2$

2. Write the formula.

a. sodium bromide b. calcium oxide c. zinc sulfide

d. silver bromide e. aluminum oxide f. sodium fluoride

g. zinc oxide h. sodium chloride i. barium fluoride

j. calcium chloride k. aluminum phosphide

l. aluminum nitride m. hydrogen sulfide n. barium hydride

o. magnesium nitride p. hydrogen selenide q. lithium chloride

r. potassium oxide s. barium phosphide t. potassium iodide

u. beryllium nitride v. calcium nitride w. aluminum chloride

x. hydrogen fluoride y. gallium oxide z. zinc arsenide

i. cadmium phosphide ii. calcium fluroide iii. californium bromide

MOLECULAR COMPOUNDS BOOKLET
SCIENCE 90

1. Write the names of the following formulas.

a. N_2O_5 b. SO_2 c. SF_6

d. OF_2 e. P_2O_4 f. IBr_3

g. NH_3 h. P_2O_5 i. HCl

j. Cl_2 k. CO l. CCl_4

m. ICl_5 n. P_4O_{10} o. O_2

p. PCl_5 q. N_2O_3 r. PF_5

s. ClO_2 t. N_2H_4 u. NBr_5

2. Give the correct formula.

a. silicon dioxide b. sulfur trioxide c. oxygen

d. diphosphorous pentaoxide e. phosphorous pentachloride

f. oxygen diflouride g. carbon tetraflouride h. nitrogen triiodide

i. nitrogen monoxide j. tetranitrogen octaoxide k. flourine

l. nitrogen dioxide m. tetrasulfur dinitrogen n. bromine chloride

o. sulfur

p. tetraphosphorous decaoxide

q. dihydrogen dioxide

r. carbon dioxide

s. nitrogen gas

t. phosphorous triiodide

Practice Problems - Both Molecular and Ionic

Part I: Name the following compounds

a. CH_4 b. KBr c. HCl d. PF_5 e. SiH_4

f. N_2H_4 g. I_2O_5 h. Al_2S_3 i. CaO j. HBr

k. Ca_3P_2 l. PF_5 m. F_2 n. NH_3 o. S_8

Part II: Write the formula for these compounds.

a. bromine b. nitrogen dioxide c. sodium phosphide

d. sodium chloride e. aluminum oxide f. barium fluoride

g. potassium selenide h. dinitrogen trioxide i. Sulfur

j. lithium fluoride k. cesium hydride l. aluminum oxide

m. sulfur octanitride n. glucose o. phosphorus pentachloride